**Molarity Worksheet Name:\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

Write the formula for molarity out in words & in units:

Problems: Show all work and circle your final answer.

1. To make a 4.00 M solution, how many moles of solute will be needed if 12.0 liters of solution are required?

2. How many moles of sucrose are dissolved in 250 mL of solution if the solution concentration is 0.150 M?

3. What is the molarity of a solution of HNO3 that contains 12.6 grams HNO3 in 1.0 L of solution?

4. How many grams of potassium nitrate are required to prepare 0.250 L of a 0.700 M solution?

5. 125 cm3 of solution contains 3.5 moles of solute. What is the molarity of the

solution?

6. Which solution is more concentrated? Solution “A” contains 50.0 g of CaCO3 in

500.0 mL of solution. Solution “B” contains 6.0 moles of H2SO4 in 4.0 L of solution. SHOW WORK!

7. How many liters of solution can be produced from 2.5 moles of solute if a 2.0 M solution is needed?

8. What would be the concentration of a solution formed when 1.00 g of NaCl are dissolved in water to make 100.0 mL of solution?